

Section 111 Describing Chemical Reactions Pages 321 329 Answer Key

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111 Describing Chemical Reactions Answers

This section explains how to write equations describing chemical reactions using appropriate symbols. It also describes how to write balanced chemical equations when given the names or formulas of the reactants and products in a chemical reaction. Writing Chemical Equations (pages 321-323) 1. A

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chemical reaction occurs when one or more ...

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321-329)

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$...

Chapter 11: Chemical Reactions

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11.1 Describing Chemical Reactions 11

Section 11.111.1. Objectives. Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction. New Vocabulary stoichiometry mole ratio

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The term oxidation was originally used to describe chemical reactions involving O_2 , but its meaning has evolved to refer to a broad and important reaction class known as oxidation-reduction (redox) reactions. A few examples of such reactions will be used to develop a clear picture of this classification.

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